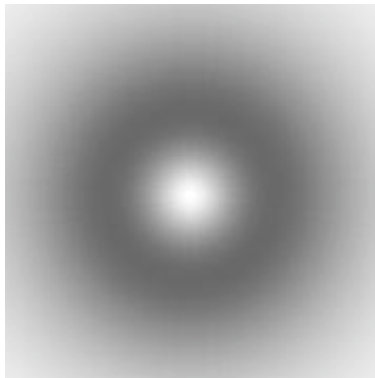


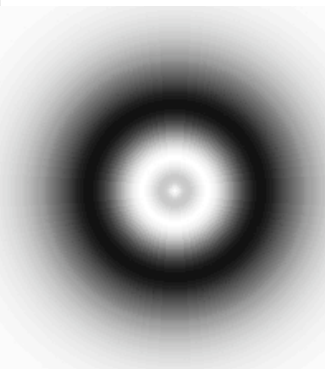
Atome d'hydrogène

Densités de probabilités de présence de l'électron :

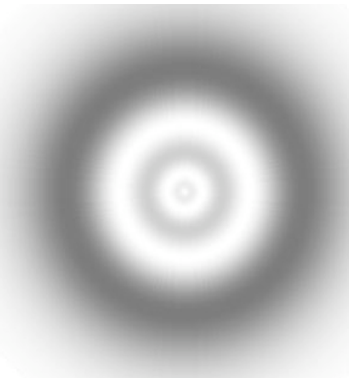
$$|\psi_{nlm}(\vec{r})|^2 = r^2 |R_{nl}(r)|^2 |Y_l^m(\theta, \varphi)|^2$$



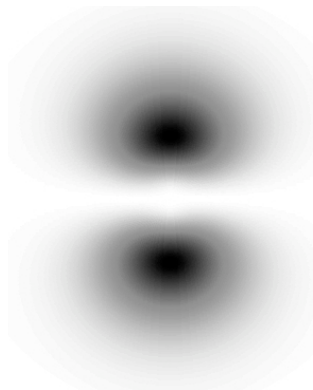
|100>
orbitale 1s



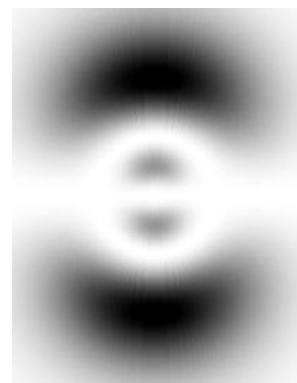
|200>
orbitale 2s



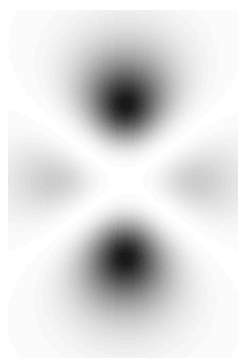
|300>
orbitale 3s



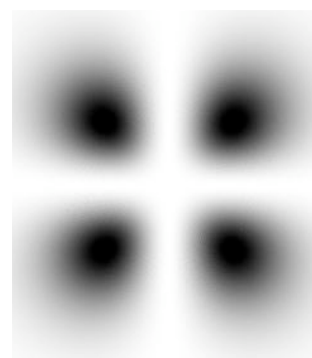
|210>
orbitale 2p



|310>
orbitale 3p



|320>
orbitale 3d



|321>
orbitale 3d

Les zones les plus sombres correspondent à la plus grande densité de probabilité de présence de l'électron.

Figures extraites de : <http://cronodon.com/Atomic/AtomTech4.html>